OCR (B) Chemistry A-Level

DM1 - Formulae, Equations and Amount of Substance

Flashcards

What is a redox titration?

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A redox titration is a titration of a reducing agent by an oxidising agent or vice versa.

What equipment is used to carry out a titration?

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- A pipette and pipette filler are used to measure out the volume of a reactant accurately before transferring it into a conical flask, where the reaction will occur.
- A burette is used as it is a controlled way to add small volumes of one reactant to another reactant.

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MnO_4^- titrations can be used to analyse a variety of reducing agents, for example:

- Fe^{2+} ions
- Ethanedioic acid $(COOH)_2$

The MnO_4^- is reduced to Mn^{2+}.

How do you do a titration with potassium manganate (VII)?

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- Add the potassium manganate(VII) solution into the burette.
- The other reactant solution is acidified with dilute H_2SO_4 and placed into a conical flask.
- The potassium manganate(VII) solution flows into the flask and as it reacts, it becomes colourless. When there is the first trace of a permanent pale pink solution, close the burette tap.